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# Energy Frequency Wavelength Worksheet Answers

**name: key period: speed /frequency / wavelength** - energy / frequency / wavelength energy (j) =  $h \times \nu$  h (planck's constant) =  $6.626 \times 10^{-34} \text{ j} \cdot \text{s}$  (joules) 10. calculate the energy of a photon of radiation with a frequency of  $8.5 \times 10^{14} \text{ hz}$ .  $-5.63 \times 10^{-19} \text{ j}$  11. calculate the energy of a gamma ray photon whose frequency is  $5.02 \times 10^{20} \text{ hz}$ ?  $3.33 \times 10^{-13} \text{ j}$  12. **wavelength, frequency & energy worksheet** - wavelength, frequency & energy worksheet . 13. how much energy does an electron lose when it emits a photon with a frequency of  $6.55 \times 10^{15} \text{ hz}$ ? 14. what is the frequency of a microwave whose wavelength is  $15.0 \text{ cm}$ ? 15. what is the frequency of a gamma ray whose wavelength is  $0.0755 \text{ nm}$ ? **wavelength, frequency, speed & energy worksheet** - wavelength, frequency, speed & energy worksheet  $c = \lambda \nu$   $\nu = c / \lambda$   $\lambda = c / \nu$   $e = h \nu$   $e = h c / \lambda$   $c = \text{speed of light } (3.0 \times 10^8 \text{ m/s})$   $\lambda = \text{wavelength}$   $\nu = \text{frequency}$   $e = \text{energy}$   $h = \text{planck's constant } (6.6262 \times 10^{-34} \text{ j}\cdot\text{s})$  1. calculate the  $\lambda$  given the  $\nu$  of radiation is  $5.10 \times 10^{14} \text{ s}^{-1}$  2. **chemistry worksheet - wavelength, frequency, & energy of ...** - chemistry worksheet and answers wavelength, frequency, & energy of electromagnetic waves.  $c = \lambda \nu$   $e = h \nu$   $c = 3.00 \times 10^8 \text{ m/s}$   $h = 6.626 \times 10^{-34} \text{ j}\cdot\text{s}$  (or  $\text{j/hz}$ ) 1. what is the wavelength of a wave having a frequency of  $3.76 \times 10^{14} \text{ s}^{-1}$ ? 2. **chemistry worksheet - wavelength, frequency, & energy of ...** - chemistry worksheet - wavelength, frequency, & energy of electromagnetic waves. show all equations, work, units, and significant figures in performing the following calculations. identify the type of radiation in each ... what is the wavelength of a wave having a frequency of  $3.76 \times 10^{14} \text{ s}^{-1}$ ? 2. **wavelength, frequency, speed & energy worksheet** - 5 25. calculate the frequency of light with wavelength =  $2.50 \times 10^{-7} \text{ m}$  26. what is the energy of cell phone radiation with a  $1 \text{ m}$  wavelength? 27. what is the energy (joules) of violet light with a frequency =  $7.50 \times 10^{14} \text{ s}^{-1}$ . 28. the formula  $\lambda \nu = h/mv$  stated that it was now possible to calculate the \_\_\_\_\_ of an electron **planck's equation name chem worksheet 5-2** - calculate the energy of photons, given their frequency. if the wavelength is given, the energy can be determined by first using the wave equation ( $c = \lambda \nu$ ) to find the frequency, then using planck's equation to calculate energy. use the equations above to answer the following questions. 1. ultraviolet radiation has a frequency of  $6.8 \times 10^{15} \text{ s}^{-1}$  ... **wavelength, frequency, speed & energy worksheet** - name \_\_\_\_\_ chapter 7 wavelength, frequency, speed & energy practice worksheet formulas and constants:  $f = c/\lambda$   $e = hf$   $c = \text{speed of light } (3.0 \times 10^8 \text{ m/s})$   $\lambda = \text{wavelength}$   $f = \text{frequency}$   $e = \text{energy}$   $h = \text{planck's constant } (6.6262 \times 10^{-34} \text{ j}\cdot\text{s})$  1. 14 calculate the  $\lambda$  given the frequency of radiation is  $5.10 \times 10^{14} \text{ hz}$ . **em spectrum, wavelength, frequency, and energy worksheet** - unit 4 - electrons & electron arrangement notes/worksheets - honors 1 em spectrum, wavelength, frequency, and energy worksheet 1.) look at the em spectrum below to answer this question. **em spectrum, wavelength, frequency, and energy worksheet** - em spectrum, wavelength, frequency, and energy worksheet 1.) look at the em spectrum below to answer this question. as you move across the visible light spectrum from red to violet... **more practice: energy, frequency, wavelength and the ...** - a photon with enough energy,  $5.1 \text{ electron volts (eV)}$  of energy - to be precise, will eject an electron from a piece of gold! what frequency and wavelength does light **sch3u7 chemistry worksheet: wavelength, frequency ...** - sch3u7 chemistry worksheet: wavelength, frequency, & energy of electromagnetic waves. show all equations, work, units, and significant figures in performing the following calculations. identify the type of radiation in each problem. (use your electromagnetic spectrum)  $c = \lambda \nu$   $e = h \nu$   $c = 3.00 \times 10^8 \text{ m/s}$   $h = 6.626 \times 10^{-34} \text{ j}\cdot\text{s}$  (or  $\text{j/hz}$ ) **electromagnetic spectrum practice problems and review ...** - wavelength, speed, frequency, energy 1. draw the electromagnetic spectrum. include all parts. label how frequency, wavelength, and energy are all related. 2. a microwave is part of the electro-magnetic spectrum therefore it's speed is the speed of light. what is its wavelength if its frequency is  $3.44 \times 10^9 \text{ hertz}$ ? 3.

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